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Scientific paper

ISSN 0351-9465, E-ISSN 2466-2585

<https://doi.org/10.5937/zasmat2203230S>Zastita Materijala 63 (3)
230 - 237 (2022)

Electrochemical studies on the corrosion behaviour of Aluminium in an alkaline medium

ABSTRACT

Corrosion of Aluminum metal at pH 10 has been controlled by dicarboxylic acids such as oxalic acid (OA) and adipic acid (AA). The Inhibition efficiency (IE) has been determined by the classical weight loss method. The maximum inhibition efficiency (IE) offered by the oxalic acid (OA) and adipic acid (AA) 250 ppm and Zn²⁺ 50 ppm system are 88 and 96%. To determine the values of Linear Polarization Resistance (LPR) and corrosion Current (I_{corr}), potentiodynamic polarization study has been used.

Keywords: Aluminium, Oxalic acid, Adipic Acid, electrochemical study.

1. INTRODUCTION

Aluminium is the second most rich metal after Iron because of its little atomic mass and negative average electrode prospective. Aluminium has a major responsibility in distinct human being activities. Because of its high energy density, it attracts the anode material for influence sources. Due to its strong electrical conductivity and good potency, it is used in construction, packing and shipping. Aluminium is commonly used by people because it is simply obtainable. It is used in numerous Industries, since it has high energy density, strength and purity[1]. Though Aluminium has adhesive shielding passivating oxide layer, this layer is amphoterically susceptible. It is establish that Aluminium metal dissolves in acidic solution and in basic solution [2-3]. Though Orsted discovered Aluminium (Latin – alumen, alum), Thanks to Wohler who isolated Aluminum [4]. Aluminium was extracted from Bauxite ore by Pierre Breathier [5]. The Principle technique used in extraction of Aluminium have been separately developed by Charles Martin Hall of Ohio in U.S and Paul Heroult of France which is commonly

referred as Hall - Heroult electrolytic procedure. Since of having 59% Copper conductivity, Aluminium behaves as a excellent thermal and electrical conductor. Aluminium is an excellent Super Conductor [6] and it is one of the mainly generally used Non Ferrous metal[7]. It is observed that global fabrication of Aluminium is very high [8,9]. A high proportion of Aluminium alloys is used for producing Aluminium foils and beverage cans[10].

Aluminium products can be used with very high effectiveness[11]. Metal and alloys decay when they come in contract with environment, especially in presence of oxygen and water. This impulsive, natural and thermodynamically favorable process is known as Corrosion [12-14]. When Aluminum is exposed to air, first it undergoes corrosion and an oxide layer is generated from prevents further Corrosion of the metal[15].

The work deals with:

1. To evaluate the inhibition efficiency (IE) of oxalic acid (OA) and adipic acid (AA) is scheming deterioration of aluminium in an sedimentary resolution incorporating 60 ppm of Cl⁻ at pH 10 in the absent and in the present of Zn²⁺ by mass (or) weight loss method.

2. To revise the mechanistic features of deterioration inhibition by electrochemical studies such seeing that polarization study and AC impedance spectra.

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Paper received: 10. 04. 2022.

Paper accepted: 12. 04. 2022.

Paper is available on the website: www.idk.org.rs/journal

2. EXPERIMENTAL SECTION

2.1. Specimens Preparation

Commercial Aluminium specimens containing 95% pure Aluminium of requisite measurement 4.0 x 1.0 x 0.2cm. The specimens were elegant to get mirror concluding. After polishing, the specimens were used in this work.

2.2. Inhibitor Solution

In 100ml of double distilled water and NaOH solution, 1g of oxalic acid and adipic acid was dissolved. The pH of this Inhibitor solution was adjusted to 10.

2.3. Weight Loss Method

Three Aluminium specimens were immersed in 100ml of an aqueous solution at pH 10 (dilute sodium hydroxide solution) for one day in the absence and presence of Inhibitors and the synergist namely Zn^{2+} . The weight of the specimens were determined using a sensitive balance Shimadzu AY62 model before immersion and after immersion for one day. The Corrosion rates were calculated. The Inhibition Efficiency(IE) was calculated from the equation:

$$\text{Inhibition Efficiency (IE)} = [(w_1 - w_2) / w_1] \times 100\%$$

Where

w_1 = Corrosion rates in the presence of the Inhibitor

w_2 = Corrosion rates in the absence of the Inhibitor

2.4. Potentiodynamic Polarization Study

Polarization study was carried out in a 660A model of CHI electrochemical work station instrument. A three electrode assembly were used. The working electrode was Aluminium, Saturated Calomel Electrode(SCE) was used as Reference Electrode and Platinum was used as the Counter Electrode. Corrosion parameters such as Tafel slopes (b_a =anodic; and b_c =cathodic) were calculated apart from Linear Polarization Resistance(LPR), corrosion Potential E_{corr} , and corrosion Current I_{corr} .

2.5. Alternating Current(AC) Impedance Measurements

In this study also, the same instrument and same set of three electrodes were used. Corrosion parameters namely, charge transfer Resistance (Rt) and double layer Capacitance(Cdl) were derived from Nyquist plots. The Impedance value was derived from Bode plots.

2.6. Electrochemicals studies and Corrosion methods

In an electrochemical studies such as potentiodynamic polarization study and AC impedance spectra is to be used as a investigate the corrosion behavior of the various metals and alloys in

different medium. These methods give highly reliable results. Polarization study has been used to study the corrosion behavior of carbon steel in geothermal fluid of low enthalpy.

Polarization study is to be usage to assess thiophenol as an inhibitor for the corrosion of a hydrogen penetration during mild steel in acidic solutions . Pasi Kangas and Mank Newman have used polarization study to evaluate the recital of duplex chromium or stainless steels in organic compound .Pasi Kangas and Rajendran *et al* have evaluated the mechanistic aspects of corrosion inhibition of carbon steel in the occurrence of 1-hydroxy ethylidene-1, 1-diphosphonic acid, Zn^{2+} and poly acrylamide by polarization curves. The electrochemical behaviours of Monel in NaCl medium and slow flowing sea water has been studied.

3. RESULTS AND DISCUSSIONS

3.1. Analysis of Weight Loss Method

Corrosion Rate (CR) of Aluminium immersed in aqueous solution containing at pH=10 for one day in the absence and presence of Inhibitor oxalic acid(OA) and Adipic Acid (AA) has been evaluated and values are given in the Tables 1 and 2.

Table 1. Corrosion rates (CR) of Aluminium immersed in aqueous solution in the absence and presence of Inhibitor system at various concentration and the Inhibition Efficiencies (IE) obtained by Weight Loss Method for the immersion period of one day

Tabela 1. Stope korozije (CR) aluminijuma potopljenog u vodeni rastvor u odsustvu i prisustvu sistema inhibitora u različitim koncentracijama i efikasnosti inhibicije (IE) dobijene metodom gubitka težine za period potapanja od jednog dana.

Cl ⁻ ppm	AA ppm	Zn ²⁺ ppm	CR mdd	IE %
60	0	0	23.48	--
60	50	0	5.16	78
60	100	0	4.69	80
60	150	0	4.46	81
60	200	0	3.99	83
60	250	0	3.28	86

Cl ⁻ ppm	OA ppm	Zn ²⁺ ppm	CR mdd	IE%
60	0	0	23.48	--
60	50	0	7.04	70
60	100	0	6.57	72
60	150	0	6.10	74
60	200	0	5.87	75
60	250	0	5.63	76

The Inhibition Efficiencies of OA and AA are also given in these Tables. The Inhibition Efficiencies of the OA Zn^{2+} and AA- Zn^{2+} systems as a function of concentration of OA and AA are shown in Fig.1. It is observed from Table 1 that OA and AA shows some Inhibition Efficiencies and Inhibition Efficiency (IE) increases when the concentration increases. For example, 50ppm OA has 70 percent and AA has 78 percent IE while 250ppm OA has 76 percent and 250ppm AA has 86 per cent. This is due to the fact that as the concentration of OA and AA increases, the protective film (probably Aluminium OA and AA complex) formed on the metal surface goes on increasing.

Table 2. Corrosion Rates (CR) of aluminium submerged in sedimentary solution incorporated 60 ppm Cl⁻ ions in the absence and presence of inhibitor scheme at different absorption and the Inhibition Efficiency (IE) attained by weight loss method (Inhibitor system: AA and OA +50 ppm of Zn^{2+} Immersion period: 1 day pH = 10)

Tabela 2. Stope korozije (CR) aluminijuma potopljenog u sedimentni rastvor inkorporirale su 60 ppm Cl⁻ jona u odsustvu i prisustvu inhibitorne šeme pri različitoj apsorpciji i efikasnosti inhibicije (IE) postignutom metodom gubitka težine (sistem inhibitora: AA i OA + 50 ppm Zn^{2+} Period potapanja: 1 dan pH = 10)

Cl ⁻ ppm	AA ppm	Zn^{2+} ppm	CR mdd	IE %
60	0	0	23.48	--
60	0	50	19.25	18
60	50	50	2.58	89
60	100	50	2.11	91
60	150	50	1.87	92
60	200	50	1.40	94
60	250	50	1.17	96

Cl ⁻ ppm	OA ppm	Zn^{2+} ppm	CR mdd	IE %
60	0	0	23.48	--
60	0	50	19.25	18
60	50	50	4.46	81
60	100	50	3.75	84
60	150	50	3.52	85
60	200	50	3.05	87
60	250	50	2.87	88

Role of Zn^{2+} is increasing the inhibitive property of Adipic acid(AA) and Oxalic Acid (OA)

The role of Zn^{2+} is increasing the inhibitive property of adipic acid (AA) and oxalic acid (OA) is given in Table 2. The adipic acid and oxalic acid shows excellent inhibitive property in the extinct of Zn^{2+} (50 ppm). For instance the mixture consisting of adipic acid and oxalic acid (250 ppm) and Zn^{2+} (50 ppm) offers 96% and 88% corrosion inhibition efficiency. This concept is known as synergistic effect. In presence of Zn^{2+} , more inhibitor molecules are transported towards the aluminium metal surface. This enhances the formation of metal inhibitor compound on the anodal sites of the metal surface, while, cathodic sites are covered by $Zn(OH)_2$. Thus anodal response and cathodal response are restricted efficiently. This statements for the synergistic effect and enhanced inhibitive efficiency of the inhibitor - Zn^{2+} . The bond between inhibitor molecule and Zn^{2+} is tough adequate to carry the inhibitor molecule to the metal surface and weedy adequate to rupture in occurrence of the metal ion, on the plane of a metal.

Role of Zn^{2+} in increasing the inhibition efficiency of the carboxylic acids

It is interesting to note that in presence of Zn^{2+} , the inhibition efficiency offered by the two carboxylic acids, namely adipic acid and oxalic acid increases. This is due to the fact, in the presence of Zn^{2+} , more inhibitor molecules are transported towards the metal surface from the bulk of the solution. More aluminium-inhibitor complexes are formed on the anodic sites of the metal. So corrosion protection increases. Thickness of the film also increases. So, corrosion inhibition efficiency increases.

3.2. Analysis of Potentiodynamic Polarization Study (pH=10)

Polarization study has been used to confirm the formation of protective film formed on the metal surface during Corrosion Inhibition process. If a protective film is formed on the metal surface, the Linear Polarization Resistance value (LPR) increases and the corrosion Current value (I_{corr}) decreases. The Potentiodynamic Polarization curves of Aluminium immersed in aqueous solution containing in the absence and presence of Inhibitors are shown in Figure.1 and Figure 2. The corrosion parameters are given in Table 3 and 4. When Aluminium was immersed in aqueous solution containing the corrosion potential was -506 mV vs SCE. When OA and AA (250ppm) and Zn^{2+} (50ppm) were added to the above system, the corrosion potential shifted to the noble side -415mV and -416 mV vs SCE. This indicates that a film is formed on the Anodic sites of the metal surface. This film controls the Anodic reaction of metal dissolution by forming Al^{3+} -OA-AA complex on the

Anodic sites of the metal surface. The formation of protective film on the metal surface is further supported by the fact that the anodic Tafel slope (b_a) slightly decreases from 221mV/decade to 146 mV/decade and 221mV/decade to 218 mV/decade. Further, the LPR value increases from 27355 ohm cm^2 to 85826ohm cm^2 and 27355ohm cm^2 to 400583 ohm cm^2 the corrosion current decreases from $20.62 \times 10^{-7} A/cm^2$ to $1.140 \times 10^{-7} A/cm^2$ and $20.62 \times 10^{-7} A/cm^2$ to $3.983 \times 10^{-7} A/cm^2$ Thus, Polarization Study confirms the formation of a Protective Film on the metal surface.

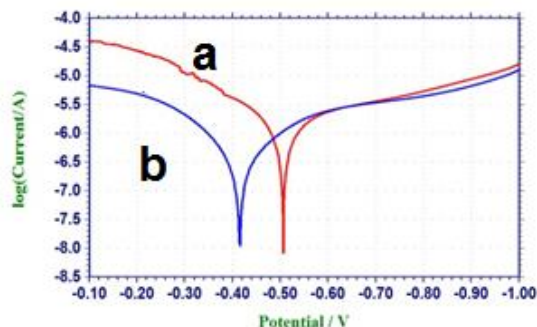


Figure 1. Polarization curves of aluminium metal submerged in experimental solutions (a)60ppmCl ions(blank)and (b) 60 ppm Cl ions + 250 ppm of OA + 50 ppm of Zn²⁺

Slika 1. Polarizacione krive metala aluminijuma potopljenog u eksperimentalne rastvore (a) 60 ppm Cl jona (prazno) i (b) 60 ppm Cl jona + 250 ppm OA + 50 ppm Zn²⁺

For blank system, the value of Current remains constant from potential 1.2 volt to 0.446 volt. At this Volt, the film is broken. Hence the corrosion Current increases suddenly. The break down potential is -0.446 volt. Similar observation is made for the Inhibitor system also. Here the break down potential is -0.38. The reason for the Current remaining constant is due to the formation of passive film on the metal surface.

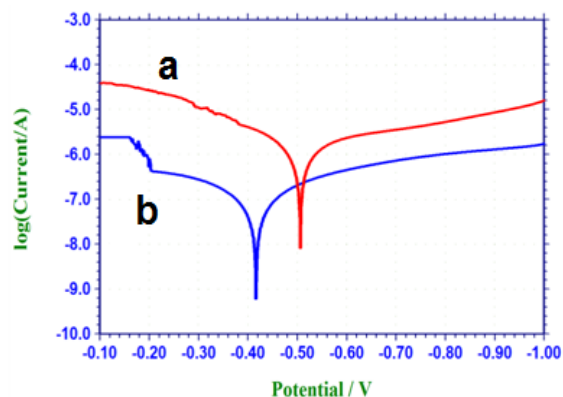


Figure 2. Polarization curves of aluminium metal submerged in experimental solutions (a) 60ppm Cl ions(blank) and (b) 60 ppm Cl ions + 250 ppm of AA + 50 ppm of Zn²⁺

Slika 2. Polarizacione krive metala aluminijuma potopljenog u eksperimentalne rastvore (a) 60 ppm Cl jona (prazno) i (b) 60 ppm Cl jona + 250 ppm AA + 50 ppm Zn²⁺

Table 3. Corrosion parameters of aluminium submerged in sedimentary solution containing 60 ppm Cl ions in the absent and present of inhibitor system obtain from potentiodynamic polarization.

Tabela 3. Parametri korozije aluminijuma potopljenog u sedimentni rastvor koji sadrži 60 ppm Cl jona u odsustvu i prisustvu inhibitornog sistema dobijeni iz potenciodinamičke polarizacije.

System	E _{corr} mV vs SCE	b _c mV/decade	b _a mV/decade	I _{corr} A/cm ²	LPR, ohm cm ²
Aqueous solution containing 60 ppm Cl ⁻ ions	-506	314	221	20.62×10^{-7}	27355
Aqueous solution containing 60 ppm Cl ⁻ ions + 250 ppm OA + 50 ppm of Zn ²⁺	-415	169	146	3.983×10^{-7}	85826

Table 4. Corrosion parameters of aluminium submerged in sedimentary solution containing 60ppm Cl ions in the absent and present of inhibitor system obtain from potentiodynamic polarization

Tabela 4. Parametri korozije aluminijuma potopljenog u sedimentni rastvor koji sadrži 60ppm Cl jona u odsustvu i prisustvu inhibitorskog sistema dobijeni iz potenciodinamičke polarizacije

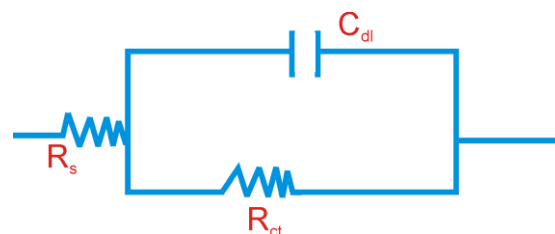
System	E _{corr} mV vs SCE	b _c mV/decade	b _a mV/decade	I _{corr} A/cm ²	LPR ohm cm ²
Aqueous solution containing 60 ppm Cl ⁻ ions	-506	314	221	20.62×10^{-7}	27355
Aqueous solution containing 60 ppm Cl ⁻ ions + AA (250 ppm) + Zn ²⁺ (50 ppm)	-416	203	218	1.140×10^{-7}	400583

3.3. Analysis of AC Impedance Spectra

AC Impedance Spectra (Electrochemical Impedance Spectra) have been used to confirm the formation of Protective Film on the metal surface [20-24]. If a protective film is formed on the metal surface, charge transfer Resistance (R_t) increases, double layer Capacitance value (C_{dl}) decreases and the Impedance $\log(z/\text{ohm})$ value increases. The AC Impedance spectra of Aluminium immersed in aqueous solution containing 60ppm Cl^- ions in the absence and the presence of Inhibitors (OA- Zn^{2+n}) and (AA- Zn^{2+}) are shown in Fig 3 and 4 (Nyquist plots) and Fig 5 and 6 (Bode plots). The AC impedance parameters namely charge transfer Resistance (R_t) and double layer Capacitance (C_{dl}) derived from Nyquist plots are given in Table 5. The Impedance $\log(z/\text{ohm})$ values derived from Bode plots are also given in Table 6.

It is ascertained that when the inhibitors OA and (250 ppm) + Zn^{2+} (50 ppm) are added, the charge transfer resistance (R_t) increases from $83.85 \Omega\text{cm}^2$ to $1072\Omega\text{cm}^2$ and $83.85\Omega\text{cm}^2$ to $2848\Omega\text{cm}^2$. The C_{dl} value decreases from $5.963 \times 10^{-9} \text{ F/cm}^2$ to $4.644 \times 10^{-9} \text{ F/cm}^2$ and $5.963 \times 10^{-9} \text{ F/cm}^2$ to $1.756 \times 10^{-9} \text{ F/cm}^2$. The impedance value

($\log(z/\text{ohm})$) increases from 1.963 to 3.280 and 1.963 to 3.622. These results lead to the conclusion that a protective film is formed on the metal surface. Equivalent circuit diagram of Aluminium immersed in various test solutions is shown in Scheme1.



R_s - Solution resistance

R_{CT} - Charge transfer resistance

C_{dl} - double layer capacitance

Scheme 1: Equivalent circuit diagram

Šema 1: Ekvivalentna šema kola

In presence of Inhibitor, the phase angle value increases from 48° to 50° . This is due to formation of Protective Film.

Table 5. Corrosion parameters of aluminium submerged in sedimentary solution containing 60 ppm Cl^- ions in the absent and present of inhibitor system obtained from AC impedance spectra (pH=10)

Tabela 5. Parametri korozije aluminijuma potopljenog u sedimentni rastvor koji sadrži 60 ppm Cl^- jona u odsustvu i prisustvu inhibitorskog sistema dobijeni iz spektra impedanse naizmjenične struje (pH=10)

System	Nyquist plot		Bode plot
	R_t ohm cm^2	C_{dl} F/cm^2	Impedance value $\log(z/\text{ohm})$
Sedimentary solution containing 60 ppm Cl^- ions	83.85	5.963×10^{-9}	1.963
Sedimentary solution containing 60 ppm Cl^- ions + 250 ppm of OA + 50 ppm of Zn^{2+}	1072	4.644×10^{-9}	3.280

Table 6. Corrosion factors of aluminium submerged in sedimentary solution containing 60 ppm Cl^- ions in the absence and presence of inhibitor system attain from AC impedance spectra (pH=10)

Tabela 6. Faktori korozije aluminijuma potopljenog u sedimentni rastvor koji sadrži 60 ppm Cl^- jona u odsustvu i prisustvu inhibitorskog sistema dobijeni iz spektra impedanse naizmjenične struje (pH=10)

System	Nyquist plot		Bode plot
	R_t ohm cm^2	C_{dl} F/cm^2	Impedance value $\log(z/\text{ohm})$
Sedimentary solution containing 60 ppm Cl^- ions	83.85	5.963×10^{-9}	1.963
Sedimentary solution containing 60 ppm Cl^- ions + AA (250 ppm) + Zn^{2+} (50 ppm)	2848	1.756×10^{-9}	3.622

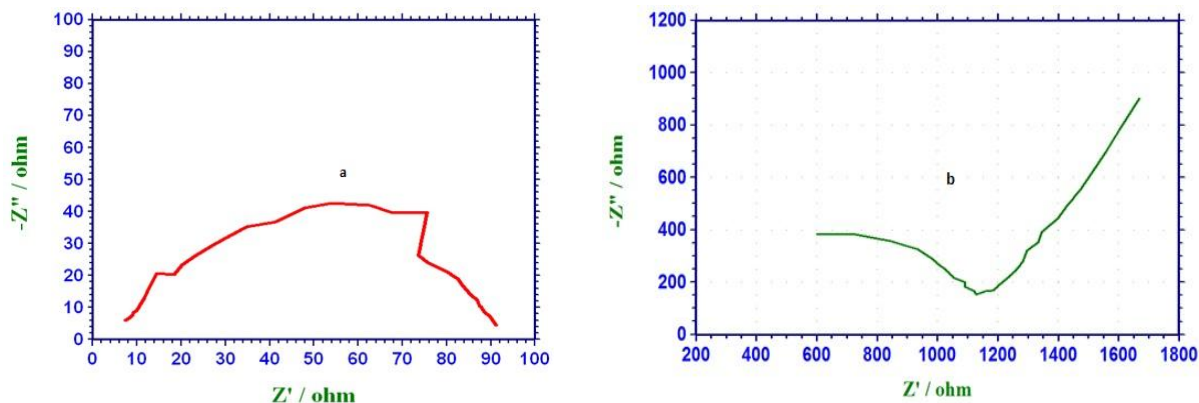


Figure 3. AC impedance spectra of aluminium metal submerged in different experimental solution (Nyquist plot) (a) 60 ppm Cl⁻ ions (blank) and (b) 60 ppm Cl⁻ ions + 250 ppm of OA + 50 ppm of Zn²⁺

Slika 3. Spektri impedanse naizmenične struje metala aluminijuma potopljenog u različiti eksperimentalni rastvor (Nyquist kriva) (a) 60 ppm Cl⁻ jona (prazno) i (b) 60 ppm Cl⁻ jona + 250 ppm OA + 50 ppm Zn²⁺

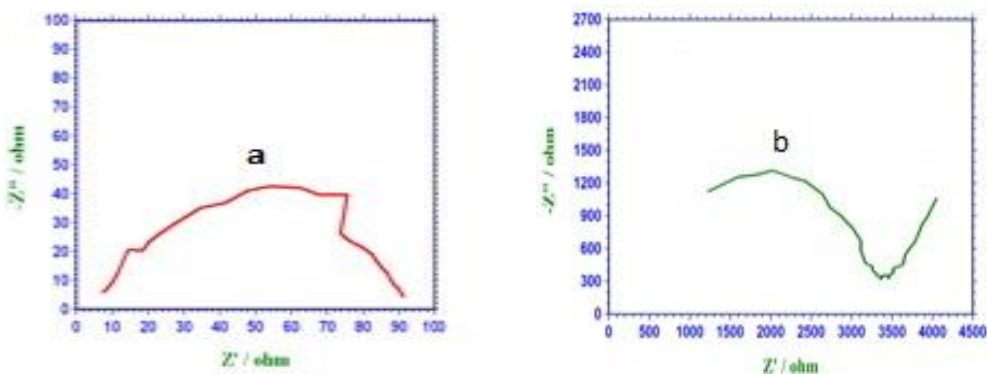


Figure 4. AC impedance spectra of aluminium metal submerged in different experimental solution (Nyquist plot) (a) 60 ppm Cl⁻ ions (blank) and (b) 60 ppm Cl⁻ ions + 250 ppm of AA + 50 ppm of Zn²⁺

Slika 3. Spektri impedanse naizmenične struje metala aluminijuma potopljenog u različiti eksperimentalni rastvor (Nyquist kriva) (a) 60 ppm Cl⁻ jona (prazno) i (b) 60 ppm Cl⁻ jona + 250 ppm AA + 50 ppm Zn²⁺

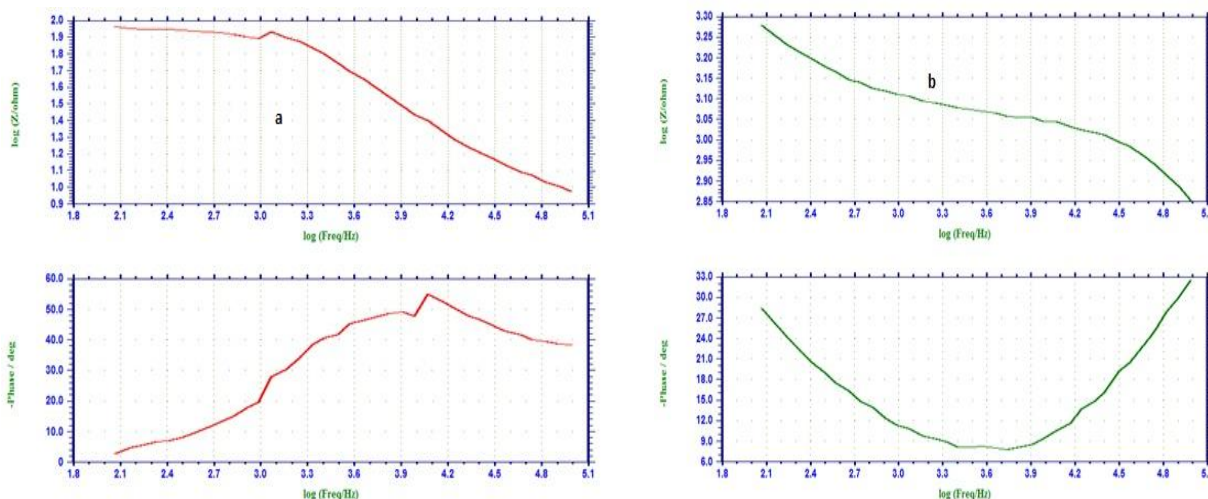


Figure 5. Bode plots of aluminum metal submerged in sedimentary solution containing (a) 60 ppm Cl⁻ ions (blank) and (b) 60 ppm Cl⁻ + 250 ppm of OA + 50 ppm of Zn²⁺

Slika 5. Bodeove krive metala aluminijuma potopljenog u sedimentni rastvor koji sadrži (a) 60 ppm Cl⁻ jona (prazno) i (b) 60 ppm Cl⁻ + 250 ppm OA + 50 ppm Zn²⁺

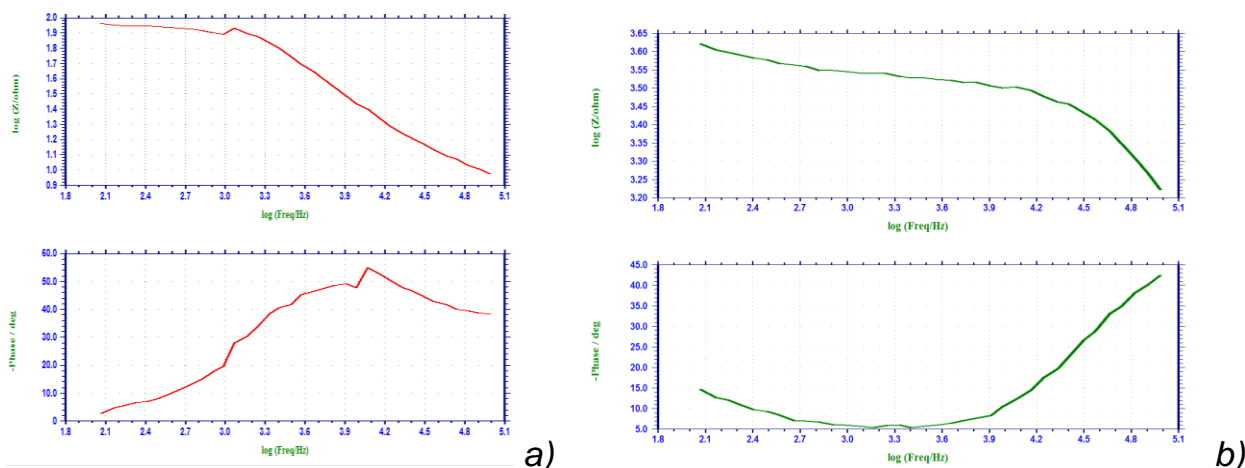


Figure 6. Bode plots of aluminium metal submerged in sedimentary solution containing (a) ppm Cl⁻ ions (blank) and (b) 60 ppm Cl⁻ + 250 ppm of AA + 50 ppm of Zn²⁺

Slika 6. Bodeovi dijagrami metala aluminijuma potopljenog u sedimentni rastvor koji sadrži (a) ppm Cl⁻ jona (prazno) i (b) 60 ppm Cl⁻ + 250 ppm AA + 50 ppm Zn²⁺

4. CONCLUSIONS

- The consequences of the weight-loss survey show that 250 ppm of AA and 50 ppm of Zn²⁺ has 96% IE in scheming corrosion of aluminium metal in an sedimentary solution incorporating 60 ppm Cl⁻ ions at pH=10.
- Polarization study expose that the preparation purpose as cathodic inhibitor calculating the cathodic reaction mainly and to some level controls the anodic reaction.
- AC impedance spectrum expose that a protecting film is formed on the metal surface.

The results of the weight-loss study show that 250 ppm of OA and 50 ppm of Zn²⁺ has 88% IE in controlling corrosion of aluminium metal in an aqueous solution containing 60 ppm Cl⁻ ions at pH=10.

- Polarization study reveals that the formulation functions as cathodic inhibitor controlling the cathodic reaction predominantly and to some extent controls the anodic reaction.
- AC impedance spectra reveal that a protective film is formed on the metal surface.

Scope for Further Study.

- It is suggested that the study of deterioration inhibition of aluminium metal with other amino acids may be carried out.
- The utmost extent of the constancy of the protecting layer created on the plane of the metal by inhibitor complex may be found.
- An exploration in the clarification of the coordination of the inhibitor compound with Al³⁺ metal ion may be done.

- An investigation in the EPR study of layer created on the metal plane or exterior may be carried out.
- Inhibition Exhibition (IE) of the preferred inhibitors may be estimated in combination with other ions such as Mn²⁺, Mg²⁺, Ba²⁺, polymers etc..

Acknowledgement

The authors are thankful to their respective management for their help and encouragement.

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IZVOD

ELEKTROHEMIJSKA ISTRAŽIVANJA O KOROZIJSKOM PONAŠANJU ALUMINIJUMA U ALKALNOJ SREDINI

Korozija metala aluminijuma na pH 10 je kontrolisana dikarboksilnim kiselinama kao što su oksalna kiselina (OA) i adipinska kiselina (AA). Efikasnost inhibicije (IE) je određena klasičnom metodom gubitka težine. Maksimalna efikasnost inhibicije (IE) koju nude sistem oksalne kiseline (OA) i adipinske kiseline (AA) 250 ppm i Zn²⁺ 50 ppm iznosi 88% i 96%. Za određivanje vrednosti linearne polarizacione otpornosti (LPR) i struje korozije (I_{corr}), korišćena je potenciodinamička studija polarizacije.

Ključne reči: aluminijum, oksalna kiselina, adipinska kiselina, elektrohemijaska studija.

Naučni rad

Rad primljen: 10. 04. 2022.

Rad prihvaćen: 12. 04. 2022.

Rad je dostupan na sajtu: www.idk.org.rs/casopis